SCIENTIFIC RESOURCES

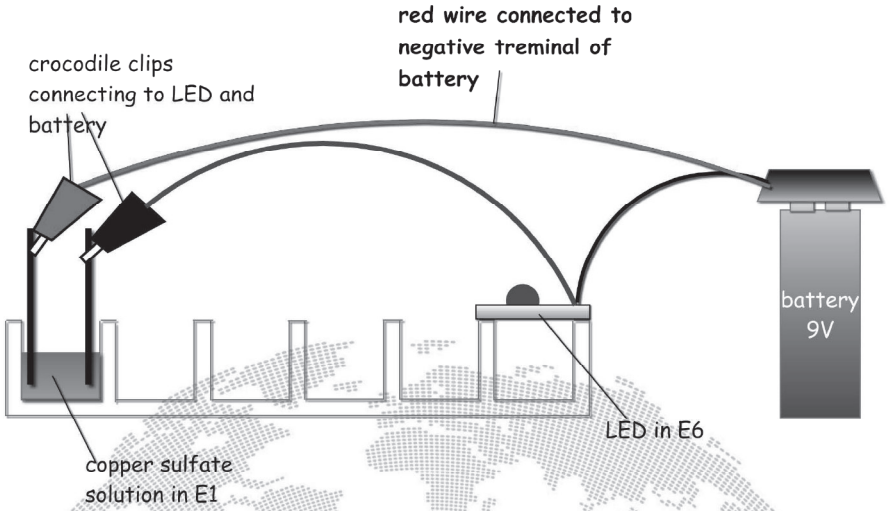
THE EXTRACTION OF COPPER FROM COPPER CARBONATE (STEP II) - THE ELECTROLYSIS OF COPPER (II) SULFATE SOLUTION

Apparatus
Comboplate®
LED + crocodile clips
9V battery
Carbon electrodes (graphite pencil leads)

Chemicals
Copper(II) sulfate solution from Step I and

Method

1. Set up the apparatus as shown in the diagram below.
2. Leave the experiment to run for a few minutes and observe what takes place at both electrodes.



Results

Electrode	Observations
Cathode (-)	
Anode (+)	

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Conclusions

Explain the observations made at both electrodes. _____

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_____

_____

Give the equation for the production of copper at the negative electrode. _____

_____

Explain why the Cu2+ ion is reduced in this reaction. _____

_____

Explain why this reaction could not take place at the anode (+). _____

_____

_____